

Chapter 4 Net Ionic Equations WS 4.1

Please write balanced net ionic equations for the following:

- | | Type of Reaction |
|---|------------------|
| 1. $\text{BaCl}_2(\text{aq}) + \text{K}_2\text{SO}_4(\text{aq}) \rightarrow$ | _____ |
| 2. $\text{Fe}_2(\text{SO}_4)_3(\text{aq}) + \text{LiOH}(\text{aq}) \rightarrow$ | _____ |
| 3. $\text{Pb}(\text{NO}_3)_2(\text{aq}) + \text{KI}(\text{aq}) \rightarrow$ | _____ |
| 4. $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{Ag}^+(\text{aq}) \rightarrow$ | _____ |
| 5. Aqueous solutions of silver nitrate and sodium phosphate are mixed. | _____ |
| 6. Aqueous solutions of calcium chloride and sodium carbonate are mixed. | |
| 7. Hydrochloric acid is neutralized by sodium hydroxide. _____ | |
| 8. $\text{Ba}(\text{OH})_2 + \text{HC}_2\text{H}_3\text{O}_2 \rightarrow$ | _____ |

9. $\text{Mg}(\text{OH})_2 + \text{HCl} \rightarrow$ _____

10. Excess chlorine gas is passed over hot iron filings. _____

11. Excess concentrated ammonia solution is added to a suspension of silver chloride. _____

12. Water is added to a sample of solid magnesium nitride. _____

13. Excess sulfur dioxide gas is bubbled through a dilute solution of potassium hydroxide. _____

14. A concentrated solution of hydrochloric acid is added to solid potassium permanganate. _____

15. Solid ammonium carbonate is heated. _____